

## 2-Hydroxybenzyl alcohol–phenanthroline (1/1)

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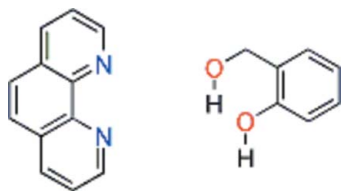
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 Key indicators: single-crystal X-ray study;  $T = 173$  K; mean  $\sigma(\text{C}–\text{C}) = 0.002$  Å;  $R$  factor = 0.038;  $wR$  factor = 0.099; data-to-parameter ratio = 13.9.

Crystals of the title compound,  $\text{C}_{12}\text{H}_8\text{N}_2 \cdot \text{C}_7\text{H}_8\text{O}_2$ , were obtained during cocrystallization experiments of a compound with two hydrogen-bond donors (2-hydroxybenzyl alcohol) with another compound containing two hydrogen-bond acceptors (phenanthroline). Unexpectedly, the two molecules do not form dimers with two  $\text{O}–\text{H} \cdots \text{N}$  hydrogen bonds connecting the two molecules. However, one of the hydroxy groups forms a bifurcated hydrogen bond to both phenanthroline N atoms, whereas the other hydroxy group forms an  $\text{O}–\text{H} \cdots \text{O}$  hydrogen bond to a symmetry-equivalent 2-hydroxybenzyl alcohol molecule. In addition, the crystal packing is stabilized by  $\pi–\pi$  interactions between the two phenanthroline ring systems, with a centroid–centroid distance of 3.570 Å.

### Related literature

For co-crystallization experiments, see: Ton & Bolte (2005); Tutughamiarso *et al.* (2009).



### Experimental

#### Crystal data

$\text{C}_{12}\text{H}_8\text{N}_2 \cdot \text{C}_7\text{H}_8\text{O}_2$   
 $M_r = 304.34$   
 Monoclinic,  $P2_1/n$   
 $a = 7.264$  (1) Å  
 $b = 20.256$  (3) Å  
 $c = 11.082$  (2) Å  
 $\beta = 109.13$  (3)°

$V = 1540.6$  (4) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.09$  mm<sup>-1</sup>  
 $T = 173$  K  
 $0.60 \times 0.50 \times 0.30$  mm

#### Data collection

Stoe IPDS II two-circle diffractometer  
 Absorption correction: none  
 20425 measured reflections

2885 independent reflections  
 2518 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.036$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.038$   
 $wR(F^2) = 0.099$   
 $S = 1.06$   
 2885 reflections

208 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.15$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.21$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D–H \cdots A$	$D–H$	$H \cdots A$	$D \cdots A$	$D–H \cdots A$
$\text{O1}–\text{H1O} \cdots \text{O2}^{\text{i}}$	0.93	1.69	2.6125 (14)	168
$\text{O2}–\text{H2O} \cdots \text{N1}^{\text{ii}}$	0.87	2.29	3.0390 (15)	144
$\text{O2}–\text{H2O} \cdots \text{N2}^{\text{ii}}$	0.87	2.15	2.8663 (14)	140

Symmetry codes: (i)  $x - \frac{1}{2}, -y + \frac{1}{2}, z - \frac{1}{2}$ ; (ii)  $x + \frac{1}{2}, -y + \frac{1}{2}, z + \frac{1}{2}$ .

Data collection: *X-AREA* (Stoe & Cie, 2001); cell refinement: *X-AREA*; data reduction: *X-AREA*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *XP* in *SHELXTL-Plus* (Sheldrick, 2008); software used to prepare material for publication: *SHELXL97* and *PLATON* (Spek, 2009).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: OM2291).

### References

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 Tutughamiarso, M., Bolte, M. & Egert, E. (2009). *Acta Cryst.* **C65**, o574–o578.

**supplementary materials**

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## 2-Hydroxybenzyl alcohol-phenanthroline (1/1)

C. Q. Ton and M. Bolte

### Comment

The aim of our research is the cocrystallization of two small organic compounds in order to examine the hydrogen bonds formed between hydrogen-bond acceptors and hydrogen-bond donors (Ton & Bolte, 2005; Tutughamiarso *et al.*, 2009). In this work, we wanted to cocrystallize phenanthroline and 2-hydroxybenzyl alcohol. However, the cocrystal, we obtained, did not show the expected AA/DD pattern, *i.e.* with two O—H $\cdots$ N hydrogen bonds connecting the two molecules to a dimer. However, one of the hydroxy groups forms a bifurcated hydrogen bonds to both phenanthroline N atoms, whereas the other hydroxy group forms a O—H $\cdots$ O hydrogen bond to a symmetry equivalent 2-hydroxybenzyl alcohol molecule. In addition, the crystal packing is stabilized by  $\pi$ - $\pi$  interactions between two phenanthroline ring systems forming a centrosymmetric dimer with a centroid $\cdots$ centroid distance of 3.570 Å. The second molecule is generated by the symmetry operation  $1 - x, -y, 1 - z$ .

### Experimental

The complex consisting of 1,10-phenanthroline and 2-hydroxybenzylalcohol was obtained by to the method of isothermal vaporization. 1,10-phenanthroline and 2-hydroxybenzylalcohol were added in an equimolar ratio (10 mmol) into a flask. Afterwards chloroform was added dropwise until the substances were completely dissolved. Then, the flask was sealed and set aside at room temperature. After two weeks crystals of the complex were obtained.

### Refinement

Hydrogen atoms were located in a difference Fourier map but those bonded to C were included in calculated positions [C—H = 0.93 - 0.99 Å] and refined as riding [ $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ ]. H atoms bonded to O were freely refined.

### Figures

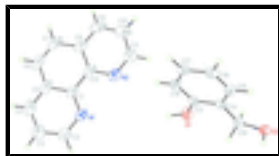


Fig. 1. A view of the molecular structure of the title compound, with the atom-numbering scheme. Displacement ellipsoids are drawn at the 50% probability level and H atoms are shown as small spheres of arbitrary radii.

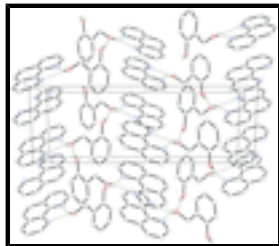


Fig. 2. Part of the crystal structure of the title compound viewed along the *c* axis. Hydrogen atoms bonded to C omitted. Hydrogen bonds shown as dashed lines.

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### Crystal data

$C_{12}H_8N_2 \cdot C_7H_8O_2$	$F_{000} = 640$
$M_r = 304.34$	$D_x = 1.312 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
Hall symbol: -P 2yn	Cell parameters from 11957 reflections
$a = 7.2640 (10) \text{ \AA}$	$\theta = 3.0\text{--}25.0^\circ$
$b = 20.256 (3) \text{ \AA}$	$\mu = 0.09 \text{ mm}^{-1}$
$c = 11.082 (2) \text{ \AA}$	$T = 173 \text{ K}$
$\beta = 109.13 (3)^\circ$	Block, colourless
$V = 1540.6 (4) \text{ \AA}^3$	$0.60 \times 0.50 \times 0.30 \text{ mm}$
$Z = 4$	

### Data collection

Stoe IPDS II two-circle diffractometer	2518 reflections with $I > 2\sigma(I)$
Radiation source: fine-focus sealed tube	$R_{\text{int}} = 0.036$
Monochromator: graphite	$\theta_{\text{max}} = 25.7^\circ$
$T = 173 \text{ K}$	$\theta_{\text{min}} = 2.8^\circ$
$\omega$ scans	$h = -8 \rightarrow 8$
Absorption correction: none	$k = -24 \rightarrow 24$
20425 measured reflections	$l = -13 \rightarrow 13$
2885 independent reflections	

### Refinement

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.038$	H-atom parameters constrained
$wR(F^2) = 0.099$	$w = 1/[\sigma^2(F_o^2) + (0.0569P)^2 + 0.2526P]$
$S = 1.06$	where $P = (F_o^2 + 2F_c^2)/3$
2885 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
208 parameters	$\Delta\rho_{\text{max}} = 0.15 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.21 \text{ e \AA}^{-3}$
	Extinction correction: none

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations

between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
N1	0.68091 (17)	0.08958 (5)	0.49625 (11)	0.0357 (3)
N2	0.89787 (17)	0.03273 (5)	0.72184 (11)	0.0371 (3)
C1	0.72912 (18)	0.02470 (6)	0.49544 (13)	0.0330 (3)
C2	0.5740 (2)	0.11589 (7)	0.38637 (14)	0.0418 (3)
H2	0.5403	0.1612	0.3860	0.050*
C3	0.5071 (2)	0.08146 (8)	0.27060 (15)	0.0478 (4)
H3	0.4303	0.1029	0.1945	0.057*
C4	0.5553 (2)	0.01591 (8)	0.26982 (15)	0.0477 (4)
H4	0.5114	-0.0087	0.1927	0.057*
C5	0.6694 (2)	-0.01440 (7)	0.38328 (14)	0.0398 (3)
C6	0.7275 (2)	-0.08270 (7)	0.38997 (17)	0.0480 (4)
H6	0.6881	-0.1088	0.3146	0.058*
C7	0.8361 (2)	-0.11014 (7)	0.50050 (18)	0.0497 (4)
H7	0.8733	-0.1552	0.5017	0.060*
C8	0.8973 (2)	-0.07275 (6)	0.61706 (15)	0.0406 (3)
C9	1.0057 (2)	-0.10017 (7)	0.73516 (18)	0.0513 (4)
H9	1.0437	-0.1452	0.7402	0.062*
C10	1.0568 (2)	-0.06217 (8)	0.84299 (18)	0.0526 (4)
H10	1.1285	-0.0803	0.9238	0.063*
C11	1.0001 (2)	0.00436 (7)	0.83078 (15)	0.0459 (4)
H11	1.0376	0.0308	0.9057	0.055*
C12	0.84467 (19)	-0.00511 (6)	0.61498 (13)	0.0342 (3)
O1	0.96385 (12)	0.19916 (5)	0.97867 (8)	0.0343 (2)
H1O	0.9040	0.1823	0.8970	0.051*
O2	1.25869 (12)	0.33737 (4)	1.24759 (8)	0.0299 (2)
H2O	1.2750	0.3699	1.2006	0.045*
C13	1.16077 (16)	0.20302 (6)	1.00360 (11)	0.0254 (3)
C14	1.26346 (16)	0.24552 (6)	1.10244 (11)	0.0252 (3)
C15	1.46410 (17)	0.25043 (6)	1.13083 (12)	0.0300 (3)
H15	1.5363	0.2785	1.1984	0.036*
C16	1.56203 (18)	0.21500 (6)	1.06219 (13)	0.0333 (3)
H16	1.6992	0.2192	1.0829	0.040*
C17	1.45820 (18)	0.17387 (6)	0.96403 (12)	0.0325 (3)
H17	1.5239	0.1501	0.9164	0.039*
C18	1.25733 (18)	0.16711 (6)	0.93473 (12)	0.0293 (3)
H18	1.1863	0.1382	0.8682	0.035*
C19	1.15095 (17)	0.28505 (6)	1.17128 (11)	0.0297 (3)
H19A	1.0333	0.3037	1.1071	0.036*

# supplementary materials

H19B                    1.1069                    0.2548                    1.2265                    0.036\*

## Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
N1	0.0415 (6)	0.0277 (5)	0.0471 (7)	-0.0037 (5)	0.0269 (5)	-0.0039 (5)
N2	0.0471 (7)	0.0279 (6)	0.0456 (7)	-0.0038 (5)	0.0277 (5)	-0.0007 (5)
C1	0.0352 (7)	0.0267 (6)	0.0486 (8)	-0.0082 (5)	0.0295 (6)	-0.0070 (5)
C2	0.0450 (8)	0.0384 (8)	0.0490 (8)	-0.0024 (6)	0.0248 (7)	0.0022 (6)
C3	0.0454 (8)	0.0580 (10)	0.0469 (8)	-0.0112 (7)	0.0244 (7)	0.0013 (7)
C4	0.0474 (8)	0.0588 (10)	0.0470 (8)	-0.0211 (7)	0.0293 (7)	-0.0152 (7)
C5	0.0415 (7)	0.0374 (7)	0.0552 (9)	-0.0164 (6)	0.0356 (7)	-0.0155 (6)
C6	0.0556 (9)	0.0368 (8)	0.0695 (11)	-0.0192 (7)	0.0449 (9)	-0.0242 (7)
C7	0.0539 (9)	0.0247 (7)	0.0892 (13)	-0.0095 (6)	0.0490 (9)	-0.0169 (7)
C8	0.0388 (7)	0.0240 (6)	0.0719 (10)	-0.0049 (5)	0.0357 (7)	-0.0035 (6)
C9	0.0459 (8)	0.0278 (7)	0.0909 (13)	0.0019 (6)	0.0371 (9)	0.0081 (8)
C10	0.0484 (9)	0.0439 (8)	0.0702 (11)	0.0004 (7)	0.0257 (8)	0.0165 (8)
C11	0.0534 (9)	0.0402 (8)	0.0509 (9)	-0.0057 (6)	0.0263 (7)	0.0027 (6)
C12	0.0365 (7)	0.0250 (6)	0.0538 (8)	-0.0066 (5)	0.0321 (6)	-0.0056 (5)
O1	0.0208 (4)	0.0510 (6)	0.0310 (5)	-0.0017 (4)	0.0083 (3)	-0.0068 (4)
O2	0.0376 (5)	0.0240 (4)	0.0281 (4)	-0.0015 (3)	0.0108 (4)	0.0035 (3)
C13	0.0224 (5)	0.0288 (6)	0.0254 (6)	0.0018 (4)	0.0084 (4)	0.0058 (5)
C14	0.0248 (6)	0.0253 (6)	0.0264 (6)	0.0022 (4)	0.0094 (5)	0.0061 (5)
C15	0.0254 (6)	0.0291 (6)	0.0349 (6)	-0.0022 (5)	0.0090 (5)	0.0024 (5)
C16	0.0232 (6)	0.0341 (7)	0.0446 (7)	0.0017 (5)	0.0138 (5)	0.0059 (6)
C17	0.0323 (6)	0.0325 (6)	0.0380 (7)	0.0079 (5)	0.0188 (5)	0.0061 (5)
C18	0.0304 (6)	0.0299 (6)	0.0283 (6)	0.0024 (5)	0.0105 (5)	0.0016 (5)
C19	0.0263 (6)	0.0326 (6)	0.0311 (6)	-0.0018 (5)	0.0104 (5)	-0.0024 (5)

## Geometric parameters ( $\text{\AA}$ , $^\circ$ )

N1—C2	1.3230 (19)	C10—C11	1.403 (2)
N1—C1	1.3609 (17)	C10—H10	0.9500
N2—C11	1.3235 (19)	C11—H11	0.9500
N2—C12	1.3562 (17)	O1—C13	1.3674 (14)
C1—C5	1.4166 (19)	O1—H10	0.9310
C1—C12	1.448 (2)	O2—C19	1.4203 (15)
C2—C3	1.400 (2)	O2—H20	0.8714
C2—H2	0.9500	C13—C18	1.3980 (17)
C3—C4	1.374 (2)	C13—C14	1.4006 (17)
C3—H3	0.9500	C14—C15	1.3901 (17)
C4—C5	1.401 (2)	C14—C19	1.5165 (16)
C4—H4	0.9500	C15—C16	1.3979 (18)
C5—C6	1.441 (2)	C15—H15	0.9500
C6—C7	1.341 (2)	C16—C17	1.3816 (19)
C6—H6	0.9500	C16—H16	0.9500
C7—C8	1.437 (2)	C17—C18	1.3938 (18)
C7—H7	0.9500	C17—H17	0.9500
C8—C9	1.402 (2)	C18—H18	0.9500

C8—C12	1.4206 (18)	C19—H19A	0.9900
C9—C10	1.366 (3)	C19—H19B	0.9900
C9—H9	0.9500		
C2—N1—C1	117.24 (12)	N2—C11—C10	124.36 (15)
C11—N2—C12	117.73 (12)	N2—C11—H11	117.8
N1—C1—C5	122.70 (13)	C10—C11—H11	117.8
N1—C1—C12	118.04 (12)	N2—C12—C8	122.19 (13)
C5—C1—C12	119.26 (12)	N2—C12—C1	118.51 (11)
N1—C2—C3	124.43 (14)	C8—C12—C1	119.30 (12)
N1—C2—H2	117.8	C13—O1—H10	110.0
C3—C2—H2	117.8	C19—O2—H2O	111.4
C4—C3—C2	118.34 (15)	O1—C13—C18	122.60 (11)
C4—C3—H3	120.8	O1—C13—C14	116.55 (10)
C2—C3—H3	120.8	C18—C13—C14	120.85 (11)
C3—C4—C5	119.67 (14)	C15—C14—C13	118.11 (11)
C3—C4—H4	120.2	C15—C14—C19	123.11 (11)
C5—C4—H4	120.2	C13—C14—C19	118.77 (10)
C4—C5—C1	117.62 (13)	C14—C15—C16	121.53 (12)
C4—C5—C6	122.97 (14)	C14—C15—H15	119.2
C1—C5—C6	119.42 (15)	C16—C15—H15	119.2
C7—C6—C5	121.31 (14)	C17—C16—C15	119.61 (11)
C7—C6—H6	119.3	C17—C16—H16	120.2
C5—C6—H6	119.3	C15—C16—H16	120.2
C6—C7—C8	121.26 (13)	C16—C17—C18	120.15 (11)
C6—C7—H7	119.4	C16—C17—H17	119.9
C8—C7—H7	119.4	C18—C17—H17	119.9
C9—C8—C12	117.59 (14)	C17—C18—C13	119.73 (12)
C9—C8—C7	122.96 (14)	C17—C18—H18	120.1
C12—C8—C7	119.44 (14)	C13—C18—H18	120.1
C10—C9—C8	120.10 (14)	O2—C19—C14	114.26 (10)
C10—C9—H9	119.9	O2—C19—H19A	108.7
C8—C9—H9	119.9	C14—C19—H19A	108.7
C9—C10—C11	118.03 (16)	O2—C19—H19B	108.7
C9—C10—H10	121.0	C14—C19—H19B	108.7
C11—C10—H10	121.0	H19A—C19—H19B	107.6
C2—N1—C1—C5	-0.08 (18)	C11—N2—C12—C1	178.66 (11)
C2—N1—C1—C12	-179.49 (11)	C9—C8—C12—N2	0.86 (18)
C1—N1—C2—C3	0.4 (2)	C7—C8—C12—N2	179.91 (11)
N1—C2—C3—C4	-0.2 (2)	C9—C8—C12—C1	-178.65 (11)
C2—C3—C4—C5	-0.3 (2)	C7—C8—C12—C1	0.40 (18)
C3—C4—C5—C1	0.55 (19)	N1—C1—C12—N2	0.52 (17)
C3—C4—C5—C6	-179.40 (13)	C5—C1—C12—N2	-178.92 (10)
N1—C1—C5—C4	-0.37 (18)	N1—C1—C12—C8	-179.96 (11)
C12—C1—C5—C4	179.04 (11)	C5—C1—C12—C8	0.61 (17)
N1—C1—C5—C6	179.58 (11)	O1—C13—C14—C15	179.69 (10)
C12—C1—C5—C6	-1.02 (17)	C18—C13—C14—C15	-0.63 (17)
C4—C5—C6—C7	-179.66 (13)	O1—C13—C14—C19	-1.68 (15)
C1—C5—C6—C7	0.40 (19)	C18—C13—C14—C19	178.01 (11)

## supplementary materials

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C5—C6—C7—C8	0.6 (2)	C13—C14—C15—C16	0.95 (17)
C6—C7—C8—C9	177.95 (13)	C19—C14—C15—C16	-177.62 (11)
C6—C7—C8—C12	-1.0 (2)	C14—C15—C16—C17	-0.24 (19)
C12—C8—C9—C10	0.1 (2)	C15—C16—C17—C18	-0.81 (18)
C7—C8—C9—C10	-178.94 (13)	C16—C17—C18—C13	1.12 (18)
C8—C9—C10—C11	-0.9 (2)	O1—C13—C18—C17	179.27 (11)
C12—N2—C11—C10	-0.1 (2)	C14—C13—C18—C17	-0.39 (17)
C9—C10—C11—N2	1.0 (2)	C15—C14—C19—O2	12.73 (16)
C11—N2—C12—C8	-0.86 (18)	C13—C14—C19—O2	-165.84 (10)

### Hydrogen-bond geometry ( $\text{\AA}$ , $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1—H1O $\cdots$ O2 <sup>i</sup>	0.93	1.69	2.6125 (14)	168
O2—H2O $\cdots$ N1 <sup>ii</sup>	0.87	2.29	3.0390 (15)	144
O2—H2O $\cdots$ N2 <sup>ii</sup>	0.87	2.15	2.8663 (14)	140

Symmetry codes: (i)  $x-1/2, -y+1/2, z-1/2$ ; (ii)  $x+1/2, -y+1/2, z+1/2$ .



Fig. 1

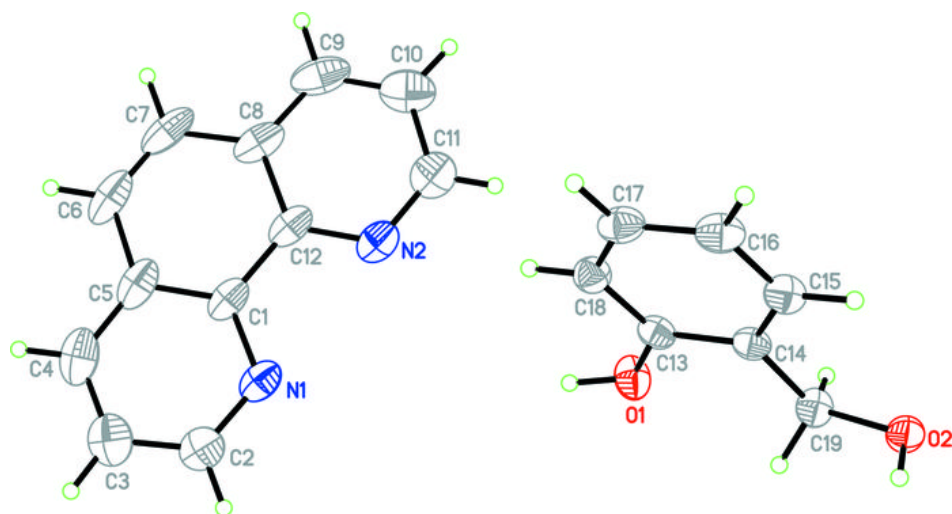


Fig. 2

